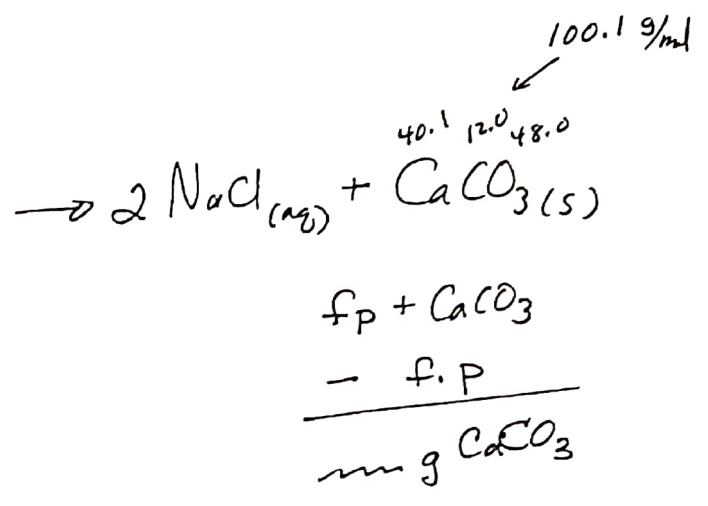
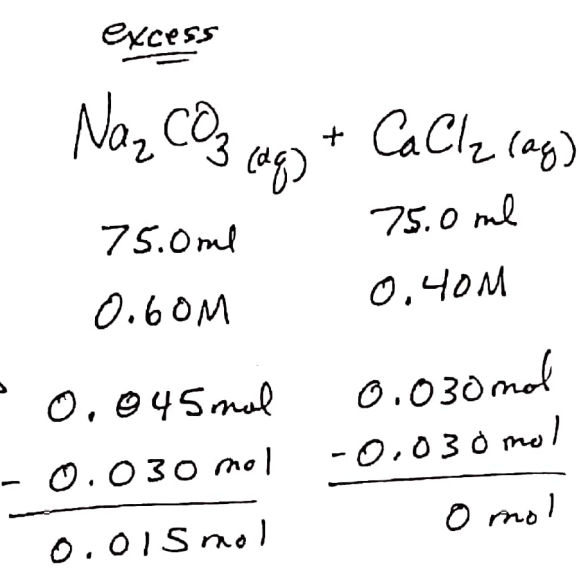


Post lab

$M = \frac{\text{mol}}{L}$
 $\text{mol} = M \times L$

Used →
 Left →



Avg: 2.79 g

$\frac{0.030 \text{ mol}}{1} \times \frac{100.1 \text{ g}}{\text{mol}} = \underline{\underline{3.0 \text{ g}}}$
 theoretical

% Yield

$\frac{2.73 \text{ g}}{3.0 \text{ g}} \times 100 = \underline{\quad} \%$