Watch Out!!
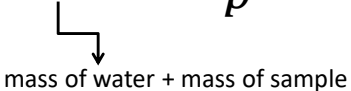
- For a tricky calorimetry problem…
  - Sample is "dissolved in water" or "reacts with water"
  - If this happens, the sample's mass must be added to the mass of water

$$q = m \, \Delta T \, C_p$$

mass of water + mass of sample

Heat of Reaction (ΔH)

- ΔH = Change in ENTHALPY

  - Enthalpy = heat absorbed/released from a chemical reaction per amount of substance

  - Units: J/g   kJ/g   J/mol   **kJ/mol** (most often)