**Made from Concentrate Exploration**

**Problem:**

How do particle pictures of solutions with various concentrations compare to one another?

**Materials:**

water

sodium chloride (NaCl - salt)

metric balance

common laboratory glassware

**Procedure:**

Label six 100 mL beakers, A through F. Measure out and add the amounts of sodium chloride (table salt) and water which are specified in the table below. Stir the contents of each beaker until the crystals dissolve. Be sure that you label each beaker as to the amount and identity of the crystals dissolved.

|  |  |  |  |
| --- | --- | --- | --- |
| **Sample** | **Volume of Water** | **Mass of Salt** | **Particle Picture** |
| A | 50 mL | 0 g |  |
| B | 50 mL | 5 g |  |
| C | 50 mL | 10 g |  |
| D | 25 mL | 0 g |  |
| E | 25 mL | 5 g |  |
| F | 25 mL | 10 g |  |

**Summing Up:**

1. Which of the solutions you prepared do you believe is the "strongest"? Why?

2. Which of the solutions you prepared do you believe is the "weakest"? Why?